

ENGINEERING DEVELOPMENT OF THE MSBR FUEL RECYCLE

REACTORS

M. E. WHATLEY, L. E. McNEESE, W. L. CARTER,
L. M. FERRIS, and E. L. NICHOLSON *Chemical Technology*
Division, Oak Ridge National Laboratory, Oak Ridge, Tennessee 37830

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The molten-salt breeder reactor being developed at Oak Ridge National Laboratory (ORNL) requires continuous chemical processing of the fuel salt, ${}^7\text{LiF}\text{-BeF}_2\text{-ThF}_4$ (72-16-12 mole%) containing ~ 0.3 mole% ${}^{233}\text{UF}_4$. The reactor and the processing plant are planned as an integral system. The main functions of the processing plant will be to isolate ${}^{233}\text{Pa}$ from the neutron flux and to remove the rare-earth fission products. The processing method being developed involves the selective chemical reduction of the various components into liquid bismuth solutions at $\sim 600^\circ\text{C}$, utilizing multistage counter-current extraction. Protactinium, which is easily separated from uranium, thorium, and the rare earths, would be trapped in the salt phase in a storage tank located between two extraction contactors and allowed to decay to ${}^{233}\text{U}$. Rare earths would be separated from thorium by a similar reductive extraction method; however, this operation will not be as simple as the protactinium isolation step because the rare-earth-thorium separation factors are only 1.3 to 3.5. The proposed process would employ electrolytic cells to simultaneously introduce reductant into the bismuth phase at the cathode and to return extracted materials to the salt phase at the anode. The practicability of the reductive extraction process depends on the successful development of salt-metal contactors, electrolytic cells, and suitable materials of construction.

INTRODUCTION

Oak Ridge National Laboratory is engaged in the development of a molten-salt breeder reactor that would operate on the ${}^{232}\text{Th}\text{-}{}^{233}\text{U}$ fuel cycle.

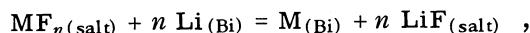
The reference reactor¹ is a single-fluid, two-region machine containing ~ 1500 ft³ of carrier salt having the composition 71.7 mole% ${}^7\text{LiF}$, 16 mole% BeF_2 , 12 mole% ThF_4 , and ~ 0.3 mole% ${}^{233}\text{UF}_4$. The reactor system would be fabricated of Hastelloy-N, and would use graphite as a moderator; corrosion of the Hastelloy is very low when $\sim 1\%$ of the uranium in the salt is present as UF_3 . Calculations have shown that single-fluid molten-salt reactors designed to operate economically at reasonable power densities and fuel inventories will not breed unless neutron absorbers such as fission products (mainly xenon and rare earths) and ${}^{233}\text{Pa}$ (which is formed from ${}^{232}\text{Th}$ and decays to ${}^{233}\text{U}$) are continually removed from the salt. Protactinium-233, which has a neutron-capture cross section of ~ 43 b, must be removed from the neutron flux on a short time cycle (3 to 5 days). Rare earths should be removed on a cycle of 30 to 60 days. The chemical processing system for effecting these separations must be close-coupled to the reactor to minimize fuel inventory.

The salt from the reactor, even after allowing 1 h for decay of very short-lived nuclides, has a specific heat generation rate of ~ 10 kW/ft³. At various places in the processing plant the protactinium and fission products will be concentrated, giving rise to heat generation rates that are 3 to 5 times this value. The protactinium isolated in the processing plant will generate ~ 5 MW of decay heat. Thus, the chemical processing system must be designed to handle much higher levels of radiation and heat generation than are encountered in the existing aqueous processes for water-cooled reactor fuels. The separations process being evaluated involves the selective reduction and extraction at $\sim 600^\circ\text{C}$ of the various species from the salt into liquid bismuth that contains lithium and thorium as the reductants. Progress in the development of this process is the subject of this paper.

CHEMISTRY OF THE REDUCTIVE EXTRACTION PROCESS

Bismuth is a noble metal that does not react with the components of the fuel salt, but will dissolve metallic lithium, uranium, thorium, and the rare earths to a reasonable extent. (Beryllium, on the other hand, is almost insoluble in bismuth.) Bismuth has a low melting point (271°C) and a high boiling point (1477°C); its vapor pressure is negligible in the temperature range of interest, 500 to 700°C. These properties, and the fact that it is almost completely immiscible with a variety of molten-fluoride salts, made it the first choice for the metal phase in the reductive extraction process.

The relative extractabilities of the important actinide and lanthanide elements were determined by measuring equilibrium distribution coefficients in the two-phase system. The extraction of a metal fluoride, MF_n , from the salt into a liquid bismuth solution can be expressed as the equilibrium reaction



in which n is the valence of the metal in the salt phase. An equilibrium constant for this reaction can be written as

$$K = \frac{a_M a_{\text{LiF}}^n}{a_{\text{MF}_n} a_{\text{Li}}^n} = \frac{X_M \gamma_M a_{\text{LiF}}^n}{X_{\text{MF}_n} \gamma_{\text{MF}_n} X_{\text{Li}}^n \gamma_{\text{Li}}^n} \quad (1)$$

in which a denotes activity, X is mole fraction, and γ is an activity coefficient. Under the experimental conditions used, a_{LiF} and the individual activity coefficients were essentially constant; thus, Eq. (1) reduces to

$$K' = \frac{X_M}{X_{\text{MF}_n} X_{\text{Li}}^n} . \quad (2)$$

The distribution coefficient for component M is defined by

$$D = \frac{\text{mole fraction of } M \text{ in bismuth phase}}{\text{mole fraction of } MF_n \text{ in salt phase}} = \frac{X_M}{X_{\text{MF}_n}} \quad (3)$$

Combination of Eqs. (2) and (3) gives

$$D = X_{\text{Li}}^n K' , \quad (4)$$

or, in logarithmic form,

$$\log D = n \log X_{\text{Li}} + \log K' . \quad (5)$$

Thus, a plot of $\log D$ vs the logarithm of the lithium concentration in the metal phase (mole fraction or at.%) should be linear with a slope equal to n . The ease with which one component can be separated from another is indicated by the ratio of

their respective distribution coefficients, i.e., by the separation factor α . If the separation factor for two components designated A and B ($\alpha = D_A/D_B$) is 1, no separation is possible; the greater the deviation of α from 1, the easier the separation.

Data²⁻⁵ obtained at 600°C using LiF-BeF₂-ThF₄ (72-16-12 mole%) as the salt phase are summarized in Fig. 1 as plots of $\log D$ vs $\log C_{\text{Li}}$. The slopes of the lines show that, under the conditions used, zirconium, thorium, and protactinium exist as tetravalent species in the salt; uranium, plutonium, and rare earths other than europium are trivalent; and only europium is reduced to the divalent state prior to extraction. Uranium and zirconium are the most easily reduced of the species shown. In fact, except for the difference in valence, their behavior is almost identical. Thus, zirconium, which is a fission product of high yield, will coextract with uranium in the reductive extraction process. Uranium and protactinium should be easily separated, and, under the proper conditions, a Pu-Pa separation is possible. Under one expected operating condition, where D_U is ~ 1 , the corresponding U-Pa and Pa-Th separation factors are ~ 100 and 3000, respectively. These separation factors comprise the basis for the protactinium isolation flow sheet. As indicated in Fig. 1, and as shown on an enlarged scale in Fig. 2, the rare-earth-thorium separation factors are only in the range of ~ 1.3 to 3.5 under the desired operating conditions ($C_{\text{Li}} > 0.1$ at.%). Thus, removal of the rare earths by the reductive extraction method will be much more difficult than the isolation of protactinium.

As noted above, most of the components of interest are adequately soluble in bismuth. Thorium is the least soluble of the extractable components; its solubility at 600°C is ~ 1800 wt ppm.⁵ Uranium and plutonium (which could be used as the fissile material for starting up an MSBR) are at least five times more soluble than thorium. Previously, no information was available on the solubility of protactinium in bismuth. However, recent work⁴ indicated its solubility to be ~ 1200 wt ppm at 500°C and >2100 ppm at 600°C. By assuming that the effect of temperature on the solubility between 500 and 700°C is about the same as it is for the other actinide metals, the solubility of protactinium at 600°C has been estimated to be 4500 wt ppm.⁴ This concentration is more than adequate to satisfy the process requirements.

Mutual solubilities of most of the major components in bismuth appear to be high enough for process application. Nickel is the only component encountered so far that causes a marked effect. The presence of as little as 100 wt ppm nickel in a bismuth solution that is nearly saturated with

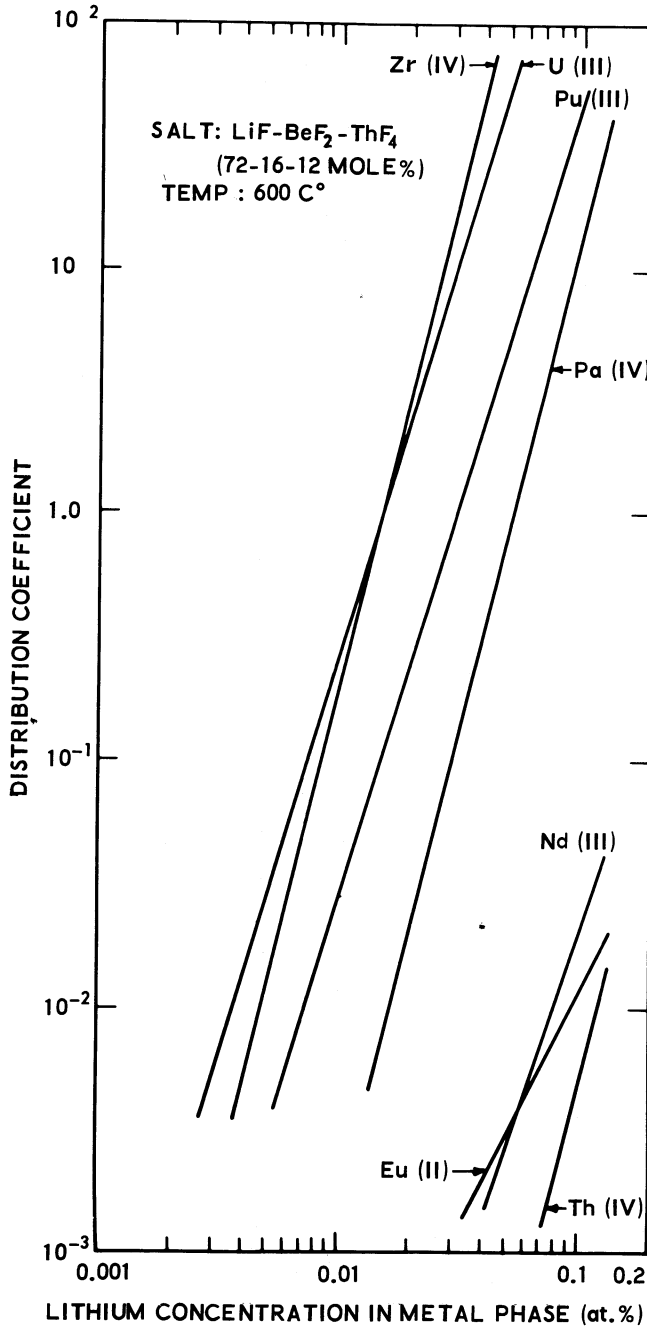


Fig. 1. Distribution coefficients of major components between a bismuth phase and a single-fluid MSBR salt.

thorium can result in the precipitation of an insoluble nickel- and thorium-containing intermetallic compound.^{2,3} One method for removing nickel is described below.

THE CONCEPTUAL PROCESS FLOW SHEET

The principal engineering features of the conceptual process^{6,7} are combined in a simplified

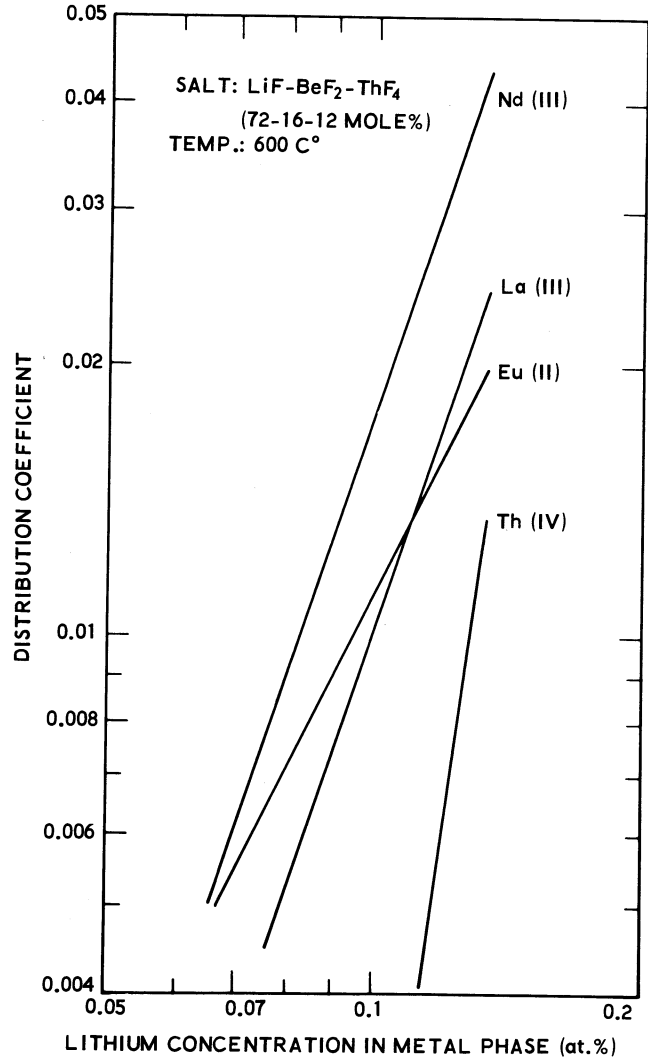
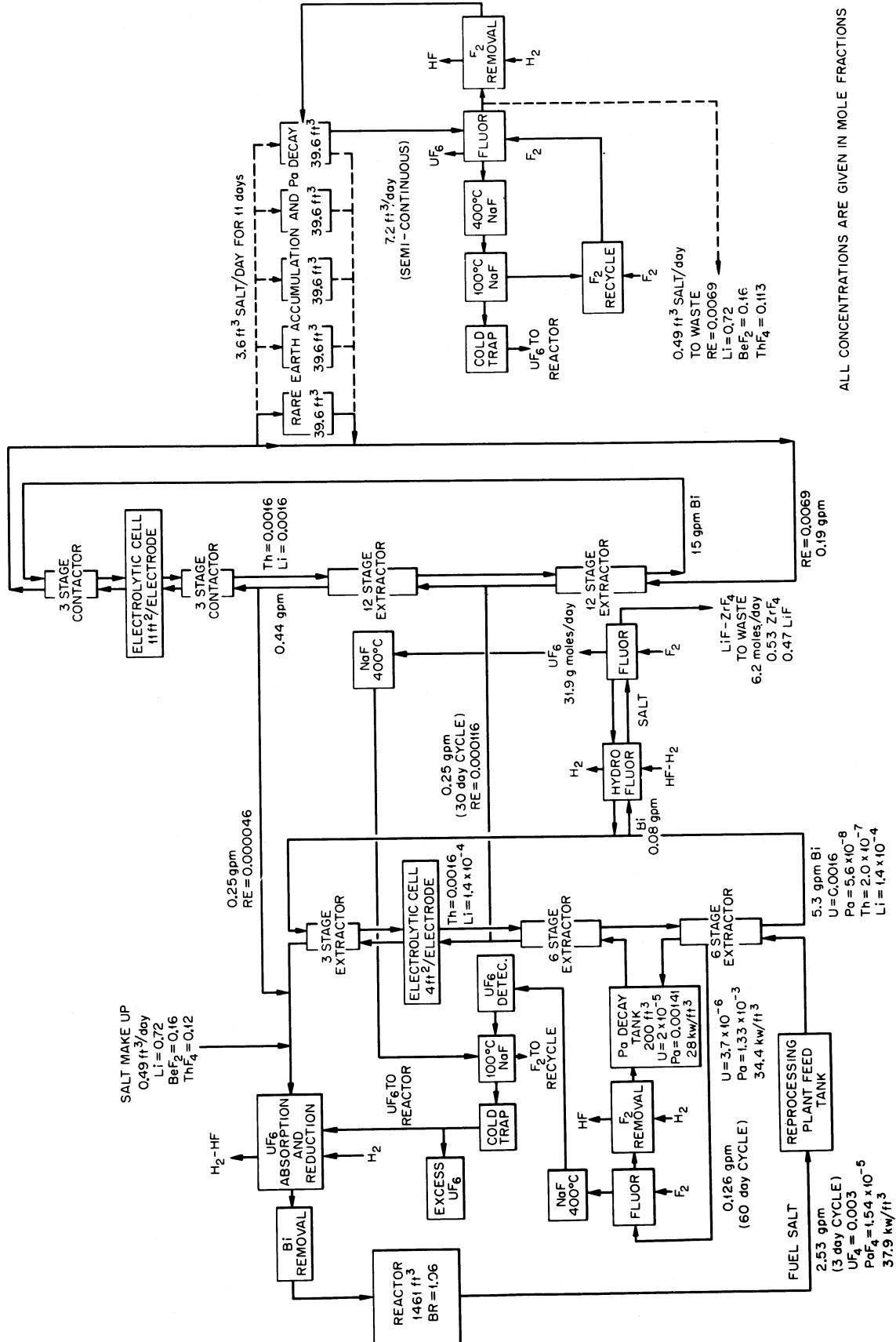


Fig. 2. Distribution of thorium and selected rare earths between a single-fluid MSBR salt and a bismuth phase.

flow sheet shown in Fig. 3. In this process fuel salt from the reactor, after <1 h of cooling, enters the bottom of the protactinium isolation system at a rate of ~2.5 gal/min. This system consists of two 7-in.-diam extractors, each having six stages. The extractors are separated by a 200-ft³ decay tank. Uranium is extracted from the fuel salt into the bismuth, and the protactinium is concentrated and trapped in the decay tank, resulting in its removal from the reactor on a 3- to 5-day cycle. The decay tank is actually a heat-exchanger in which protactinium decay heat is removed.

The bismuth is continuously circulated through the protactinium isolation system contactors and an electrolytic cell. At the anode of the cell, uranium present in the bismuth is oxidized to U⁴⁺,



ALL CONCENTRATIONS ARE GIVEN IN MOLE FRACTIONS

Fig. 3. Conceptual reductive extraction flow sheet for processing a MSBR.

which transfers to the salt stream that is returned to the reactor. At the cathode of the cell, Th^{4+} and Li^+ from the salt are reduced to metals which dissolve in the bismuth. The resulting Th-Li-Bi solution flows into the top of the upper extractor. A small side stream of salt from the lower extractor is fluorinated to remove uranium as UF_6 for control purposes. The fluorination also removes iodine, bromine, and oxygen from the salt. After treatment to remove traces of fluorine, the salt is returned to the decay tank. The UF_6 is decontaminated from fission products by passage through hot sodium fluoride beds and is collected after subsequent passage through a suitable concentration monitor. Most of the UF_6 is absorbed in salt, reduced with hydrogen to UF_4 , and returned to the reactor. Excess uranium is removed and sold.

Batch fluorination of molten salt for uranium recovery and decontamination is well-established technology. A similar operation was carried out⁸ at the MSRE when the ^{235}U fuel was replaced with ^{233}U . Small-scale tests have shown that continuous fluorination will be feasible,⁹ and that soluble UF_4 is produced when UF_6 is sorbed in salt in the presence of hydrogen.^{10,11} In the preceding operations, container corrosion could be severe; hence, protection of the wall by a layer of frozen salt is being considered.

The bismuth stream from the first stage of the lower extractor carries the uranium to the electrolytic cell in which the uranium is oxidized. About 1% of this stream is continuously treated with hydrogen fluoride in the presence of a salt to remove fission products (Zr, Zn, Ga, Cd, and Sn) and corrosion products (Fe, Ni, and Cr). After fluorination to remove the uranium, the salt is discarded to waste. This operation removes fission product zirconium on a 200-day cycle. A shorter cycle time may be necessary if nickel or other contaminants build up excessively. Hydrofluorination of this side stream of bismuth would provide a method for removing plutonium from the circuit of a molten-salt breeder reactor that was started up with plutonium.

The salt stream leaving the protactinium isolation system contains only traces of protactinium and uranium but contains practically all of the rare earths. A portion of this salt stream is withdrawn and sent to a reductive extraction process⁷ for removing rare earths. The rare-earth extraction system differs from the protactinium isolation system in that the highest concentration of rare earths occurs at the lower end of the contactor rather than in the middle. The salt-feed stream would enter near the middle of the contactor. Calculations have shown that a contactor having 24 theoretical extraction stages

with a bismuth-to-salt flow rate ratio of 80 would result in a discard salt with the rare earths ~60 times as concentrated as they occur in the reactor salt. The salt discard rate is set so that the rare earths are effectively removed on a 50-day cycle. At this discard rate, the neutron loss to rare earths in the reactor is kept at an acceptably low level, and the alkali metal and alkaline earth fission products (which remain in the salt throughout the process) are removed from the reactor on an 8- to 10-year cycle. The salt that is discarded would have a heat generation rate of ~17 kW/ft³ and would have to be stored for radioactive decay.

The reconstituted fuel salt will contain a small but unknown amount of bismuth. Most of this bismuth must be removed from the salt to ensure that its concentration in the salt returning to the reactor will not be high enough to cause corrosion of the Hastelloy.

Not all of the fission products having significant neutron capture cross sections would be removed by the reductive extraction process. Xenon, krypton, and tritium will be removed from the reactor as gases on about a 1-min cycle by a helium purge. Experience with the MSRE has shown that the noble metal fission products (e.g. Mo, Ru, Tc, Rh, Nb, and Pd) are not present in the salt as fluorides.¹² Instead, they apparently exist in the metallic state because of the reducing condition in the reactor. A portion of these metals is deposited on the surfaces of the graphite and the Hastelloy, and the rest is present in the gas phase in the form of a smoke. It is estimated that at least half of the noble metal fission products will also be removed from the MSBR by the helium purge. Thus, the reactor off-gas system must be designed to handle a significant amount of gaseous and particulate fission products.

PROTACTINIUM ISOLATION SYSTEM CALCULATIONS

The distribution coefficient data for the components of interest (Fig. 1) provide a firm basis for calculation of the separations attainable in a multistage countercurrent extraction system. A typical set of concentration profiles for this system⁷ with six theoretical stages below the decay tank and six theoretical stages above, is shown in Fig. 4. The points beyond the left margin of the figure represent the composition of the reactor fuel salt entering the extraction system. The maximum concentrations of uranium, protactinium, and thorium in the metal phase are limited by the solubility of thorium in bismuth, which, in turn, governs the salt-bismuth flow rate ratio. The protactinium concentrations in both the salt phase and the metal phase reach maxima in the vicinity of the decay tank, where the protactinium